

Name: \_\_\_\_\_

Date: \_\_\_\_\_

**BLM 3-10**

## **Teacher Demonstration: A Catalyst in a Chemical Change**

Answer these questions as your teacher prepares and completes the demonstration on page 70.

- 1.** What question will this experiment try to answer?

\_\_\_\_\_

- 2.** What reactants are being used?

\_\_\_\_\_ and \_\_\_\_\_

- 3.** What is the catalyst for this reaction?

\_\_\_\_\_

### **What Do You Think Will Happen?**

- 4.** Do you think there will be a difference in the rate of change when the catalyst is added?      YES      NO      Explain. \_\_\_\_\_

\_\_\_\_\_

- 5.** What variable is being tested? \_\_\_\_\_

- 6.** List at least three variables that are important to keep the same.

\_\_\_\_\_

\_\_\_\_\_

- 7.** On the back of the page, sketch the set-up for this demonstration. After the reactions, use shading or colour to show any colour changes you observed.

### **What Did You Observe?**

- 8.** Before any reaction takes place, observe the solutions.

**a)** Describe the Rochelle salt solution. \_\_\_\_\_

\_\_\_\_\_

**b)** Describe the hydrogen peroxide. \_\_\_\_\_

\_\_\_\_\_

**c)** Describe the cobalt chloride solution. \_\_\_\_\_

\_\_\_\_\_

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(continued)

- 9.** Describe the reaction in beaker A between the hydrogen peroxide and Rochelle salt solution.

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- 10.** Describe the reaction in beaker B after the cobalt chloride solution is added.

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- 11.** List two observations that tell you when the reaction stopped.

**a)** \_\_\_\_\_

**b)** \_\_\_\_\_

- 12.** Describe what happened when the solution in beaker B was poured into beaker A.

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### What Did You Learn?

- 13.** How did the catalyst affect the reaction?

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- 14. a)** Was there any catalyst left in the mixture?      YES      NO

**b)** How do you know? \_\_\_\_\_

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- 15.** How does a catalyst affect the rate of change?

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