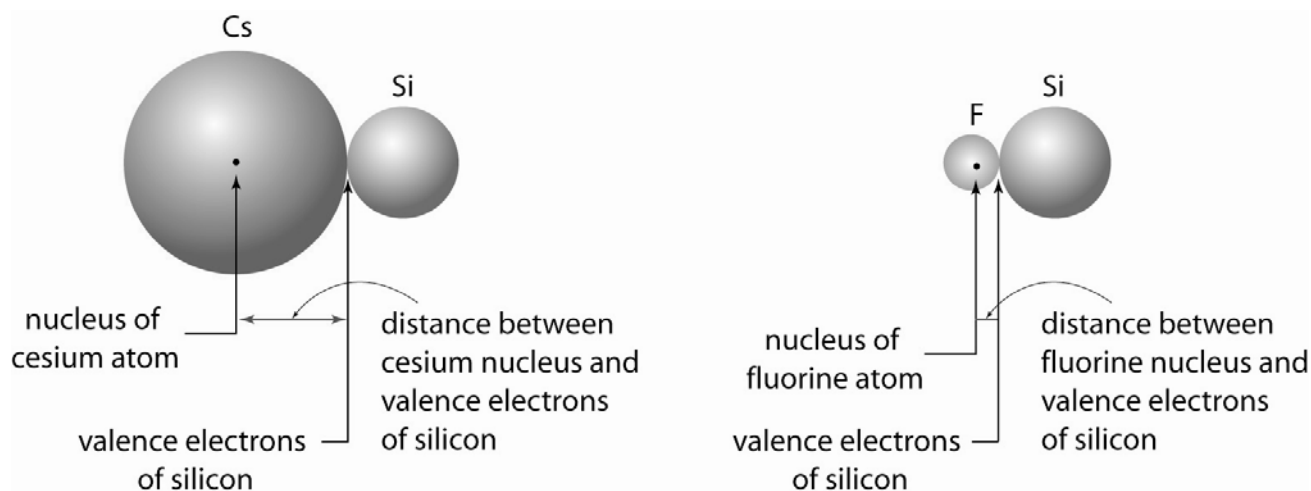


Ion Sizes and Attractive Forces



The valence electrons of silicon can get much closer to the nucleus of fluorine than to the nucleus of cesium. This is because the attractive force of the fluorine nucleus on the valence electrons of silicon is much greater than the attractive force of the cesium nucleus on the valence electrons of silicon.