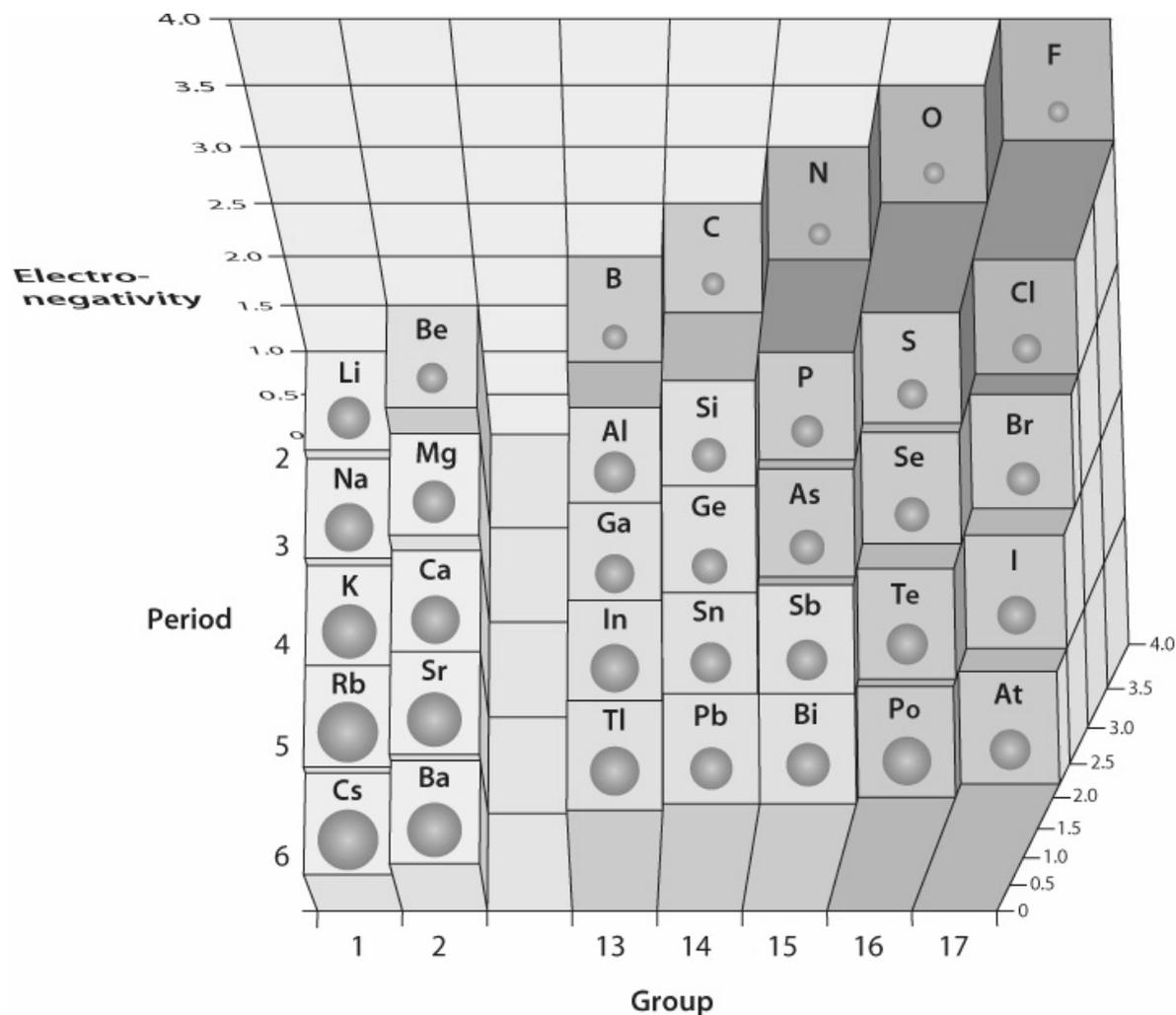


3-D Periodic Table Trends in Electronegativities and Atomic Sizes



In this three-dimensional image of the main group elements, the electronegativity scale is vertical. The period and group numbers are on the left and bottom, respectively.

Note that the trends for atomic size and electronegativity are opposite of each other. As atomic radius decreases, moving left to right in a period, the electronegativity increases. As the radius of the atoms decrease in a period, the valence electrons are found just a bit closer to the nucleus; therefore, the atom has more attraction to its own nucleus and to another atom's nucleus if bonded.