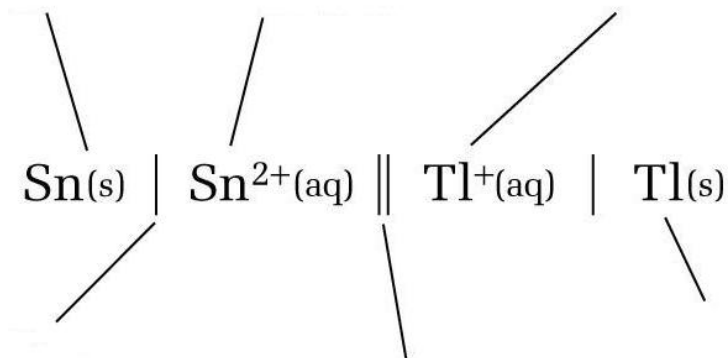


CHAPTER 13	Voltaic Cell Notation	BLM 13.1.5
ASSESSMENT		

Answer the following questions in the space provided.

1. (a) Label each component of the following shorthand representation of a galvanic cell.



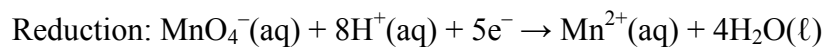
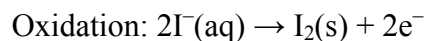
- (b) What does the location (left or right of the double vertical line) of each substance in the shorthand representation tell you about the cell?

- (c) If you were to draw a diagram of the cell, would it matter whether the cathode was shown to the right or to the left of the anode? Explain your answer.

2. Use galvanic cell notation to represent a galvanic cell that has a copper electrode in copper(II) sulfate solution and an iron electrode in iron(II) nitrate solution.

CHAPTER 13	Voltaic Cell Notation (continued)	BLM 13.1.5
ASSESSMENT		

3. Consider the following half-reactions:



(a) Write the overall reaction.

(b) A galvanic cell based on this reaction uses inert electrodes, such as graphite electrodes. Explain why.

(c) Write the shorthand representation of the cell using galvanic cell notation.

4. Use the following shorthand representation to sketch a possible design for the following cell:

