

CHAPTER 13	Electrolytic Cells Answer Key	BLM 13.3.1A
ANSWER KEY		

Voltaic cell	Electrolytic cell
Oxidation half-reaction: $\text{Zn(s)} \rightarrow \text{Zn}^{2+}(\text{aq}) + 2\text{e}^{-}$	Oxidation half-reaction: $\text{Cu(s)} \rightarrow \text{Cu}^{2+}(\text{aq}) + 2\text{e}^{-}$
Reduction half-reaction: $\text{Cu}^{2+}(\text{aq}) + 2\text{e}^{-} \rightarrow \text{Cu(s)}$	Reduction half-reaction: $\text{Zn}^{2+}(\text{aq}) + 2\text{e}^{-} \rightarrow \text{Zn(s)}$
Cell reaction: $\text{Zn(s)} + \text{CuSO}_4(\text{aq}) \rightarrow \text{ZnSO}_4(\text{aq}) + \text{Cu(s)}$	Cell reaction: $\text{Cu(s)} + \text{ZnSO}_4(\text{aq}) \rightarrow \text{CuSO}_4(\text{aq}) + \text{Zn(s)}$