


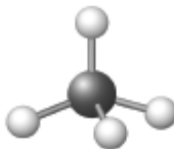
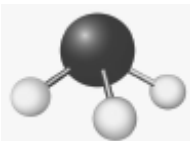
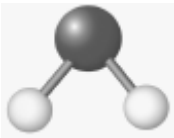


CHAPTER 2	Molecular Shapes	BLM 2.1.9
OVERHEAD		

VSEPR Class	Name of molecular shape	Type of electron pairs	Shape	Example
AX_2	linear	all BP		CO_2
AX_3	trigonal planar	all BP		CH_2O
AX_2E	bent (trigonal planar electron groups)	2 BP, 1 LP		SO_2
AX_4	tetrahedral	all BP		CH_4
AX_3E	trigonal pyramidal (tetrahedral electron groups)	3 BP, 1 LP		NH_3
AX_2E_2	bent (tetrahedral electron groups)	2 BP, 2 LP		H_2O