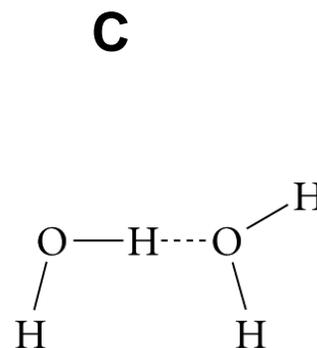
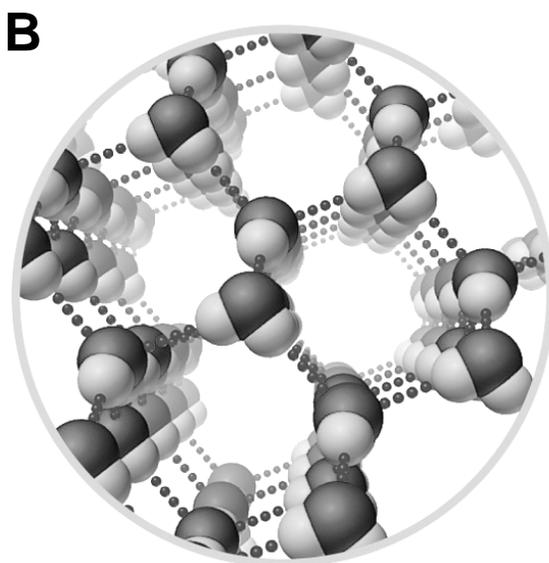


Individual hydrogen bonds are much weaker than covalent bonds. However, when many hydrogen bonds are acting on each water molecule at the same time, they have a significant effect on the properties of water.



The hydrogen bonds in ice are longer than they are in water, making ice less dense than liquid water. It is the orientation of the hydrogen bonds in ice that gives them their maximum strength.