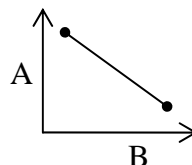


CHAPTER 3	What Do You Know About Gases?	BLM 3.0.5
ASSESSMENT		

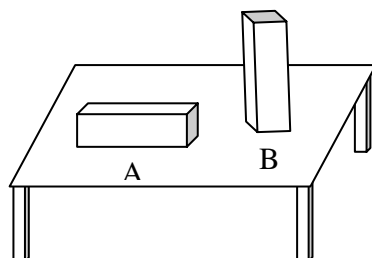
Read each statement and decide if it is true (T) or false (F). Write NS if you are not sure.

Before Unit**After unit completion**

1. The volume of a gas expands when heated.
2. At constant temperature and pressure, the volume of one mole of oxygen is greater than the volume of one mole of helium.
3. At 25 °C, the average kinetic energy of oxygen and helium are equal.
4. When a balloon is compressed, the molecules get smaller.
5. The density of liquid water is greater than the density of gaseous water.
6. A gas molecule travels in straight lines until it collides with another gas molecule or the walls of the container.
7. Substances that mix completely with each other are homogenized.
8. The volume of a tire on a loaded truck decreases because air is pushed out.
9. If two variables, A and B, exhibit an inverse relationship, the graph of the data will appear as follows:



10. If you wish to examine how volume changes with temperature in an experiment while holding pressure constant, temperature will be the responding variable.
11. Brick A exerts more pressure on the table than brick B.



12. If the pressure of a gas is increased from 50 kPa to 100 kPa, and the temperature rises from 100 °C to 200 °C, the volume will not change.
13. Four litres of oxygen combine with 8.00 L of nitrogen to form 6.00 L of mixed gases.