

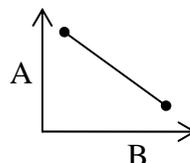
CHAPTER 3	What Do You Know About Gases?	BLM 3.0.5
ASSESSMENT		

Read each statement and decide if it is true (T) or false (F). Write NS if you are not sure.

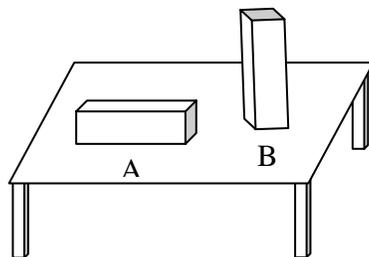
### Before Unit

### After unit completion

- The volume of a gas expands when heated.
- At constant temperature and pressure, the volume of one mole of oxygen is greater than the volume of one mole of helium.
- At 25 °C, the average kinetic energy of oxygen and helium are equal.
- When a balloon is compressed, the molecules get smaller.
- The density of liquid water is greater than the density of gaseous water.
- A gas molecule travels in straight lines until it collides with another gas molecule or the walls of the container.
- Substances that mix completely with each other are homogenized.
- The volume of a tire on a loaded truck decreases because air is pushed out.
- If two variables, A and B, exhibit an inverse relationship, the graph of the data will appear as follows:



- If you wish to examine how volume changes with temperature in an experiment while holding pressure constant, temperature will be the responding variable.
- Brick A exerts more pressure on the table than brick B.



- If the pressure of a gas is increased from 50 kPa to 100 kPa, and the temperature rises from 100 °C to 200 °C, the volume will not change.
- Four litres of oxygen combine with 8.00 L of nitrogen to form 6.00 L of mixed gases.