

Common Strong Bases:

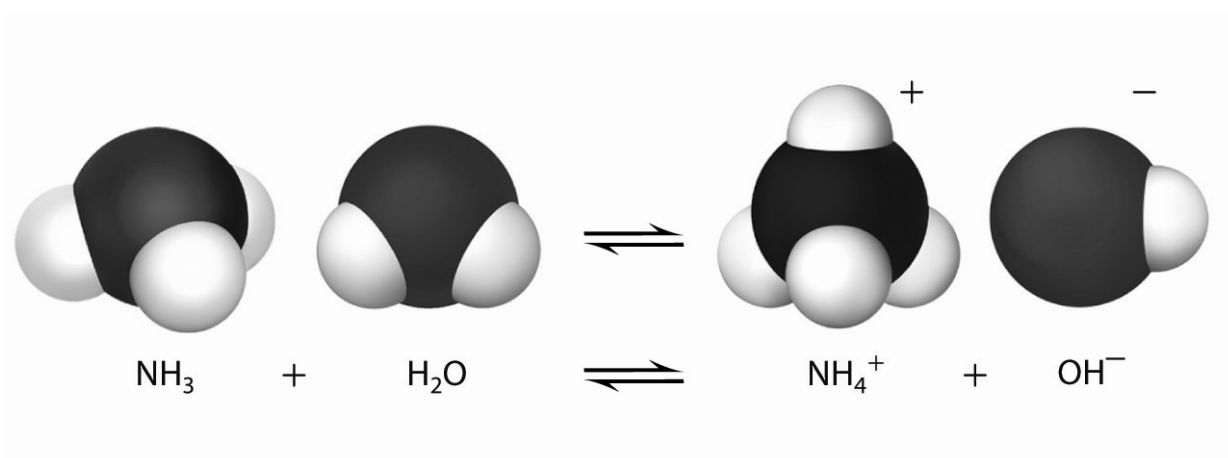
Aqueous sodium hydroxide, NaOH(aq)

Aqueous potassium hydroxide, KOH(aq)

Aqueous calcium hydroxide, $\text{Ca(OH)}_2\text{(aq)}$

Aqueous strontium hydroxide, $\text{Sr(OH)}_2\text{(aq)}$

Aqueous barium hydroxide, $\text{Ba(OH)}_2\text{(aq)}$

Weak Bases in Aqueous Solution:

About 1% of ammonia molecules react in water to produce $\text{OH}^-\text{(aq)}$. In a 0.10 mol/L solution of ammonia, the concentration of $\text{OH}^-\text{(aq)}$ is only about 0.0010 mol/L.