

The Strong Acids:

perchloric acid (aqueous hydrogen perchlorate), $\text{HClO}_4(\text{aq})$

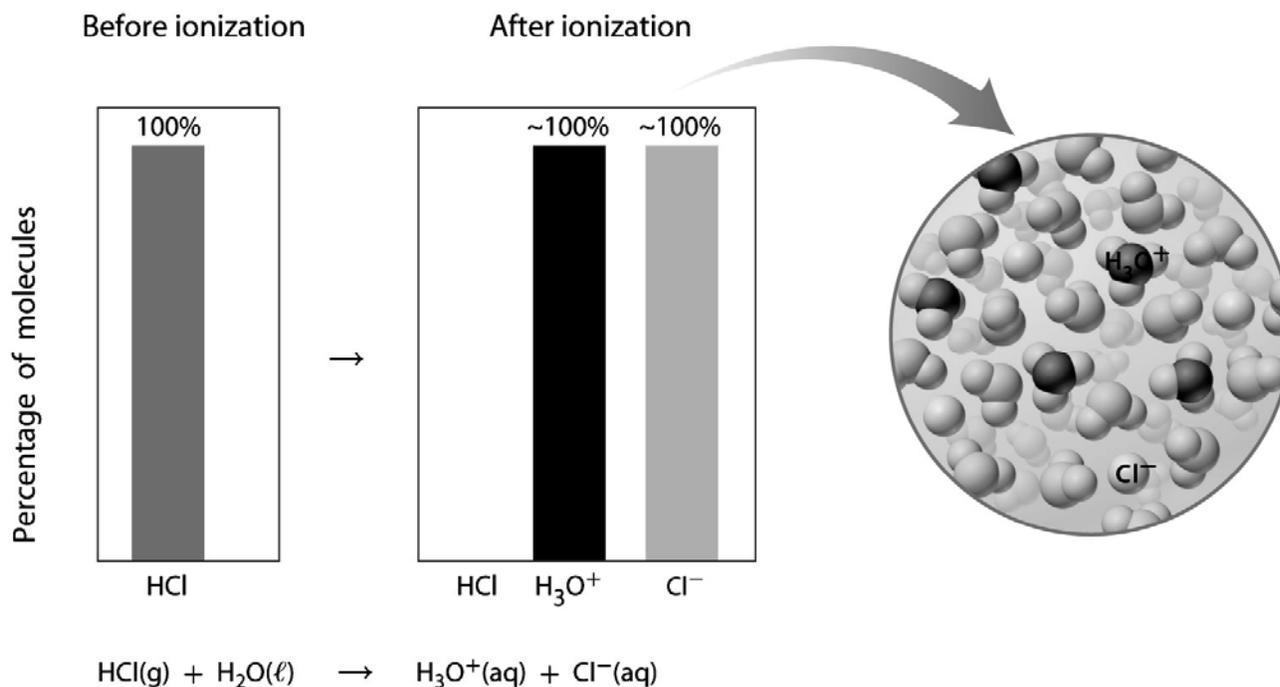
hydroiodic acid (aqueous hydrogen iodide), $\text{HI}(\text{aq})$

hydrobromic acid (aqueous hydrogen bromide), $\text{HBr}(\text{aq})$

hydrochloric acid (aqueous hydrogen chloride), $\text{HCl}(\text{aq})$

sulfuric acid (aqueous hydrogen sulfate), $\text{H}_2\text{SO}_4(\text{aq})$

nitric acid (aqueous hydrogen nitrate), $\text{HNO}_3(\text{aq})$



In an aqueous solution, close to 100% of hydrogen chloride molecules ionize to form $\text{H}_3\text{O}^+(\text{aq})$ and $\text{Cl}^-(\text{aq})$.