

CHAPTER 6	Naming Acids and Bases Review Answer Key	BLM 6.1.2A
ANSWER KEY		

1.

Name of pure substance	IUPAC name for acid	Classical name for acid	Chemical Formula
hydrogen cyanide	<i>aqueous hydrogen cyanide</i>	<i>hydrocyanic acid</i>	<i>HCN(aq)</i>
<i>hydrogen bromide</i>	<i>aqueous hydrogen bromide</i>	<i>hydrobromic acid</i>	<i>HBr(aq)</i>
<i>hydrogen sulfate</i>	<i>aqueous hydrogen sulfate</i>	<i>sulfuric acid</i>	<i>H<sub>2</sub>SO<sub>4</sub>(aq)</i>
<i>hydrogen nitrite</i>	<i>aqueous hydrogen nitrite</i>	<i>nitrous acid</i>	<i>HNO<sub>2</sub>(aq)</i>
<i>hydrogen thiocyanate</i>	<i>aqueous hydrogen thiocyanate</i>	<i>thiocyanic acid</i>	<i>HSCN(aq)</i>
hydrogen phosphate	<i>aqueous hydrogen phosphate</i>	<i>phosphoric acid</i>	<i>H<sub>3</sub>PO<sub>4</sub>(aq)</i>
<i>hydrogen chlorite</i>	<i>aqueous hydrogen chlorite</i>	<i>chlorous acid</i>	<i>HClO<sub>2</sub>(aq)</i>
<i>hydrogen iodide</i>	<i>aqueous hydrogen iodide</i>	<i>hydroiodic acid</i>	<i>HI(aq)</i>
<i>hydrogen borate</i>	<i>aqueous hydrogen borate</i>	<i>boric acid</i>	<i>H<sub>3</sub>BO<sub>3</sub>(aq)</i>

2.

Ending of name of pure substance	Classical name of acid
-ate	<i>-ic acid</i>
<i>hydrogen ----ide</i>	<i>hydro-----ic acid</i>
<i>-ite</i>	<i>-ous acid</i>

3.

(a) NaOH	<i>sodium hydroxide</i>	(c) strontium hydroxide	<i>Sr(OH)<sub>2</sub>(aq)</i>
(b) Mg(OH) <sub>2</sub>	<i>magnesium hydroxide</i>	(d) potassium hydroxide	<i>KOH(aq)</i>
(c) NH <sub>3</sub>	<i>ammonia</i>	(e) cesium hydroxide	<i>CsOH(aq)</i>

(Answers appear in italics.)