

CHAPTER 7	Predicting and Balancing Formation, Decomposition, and Hydrocarbon Combustion Reactions Answer Key	BLM 7.0.5A
ANSWER KEY		

- $2\text{C(s)} + \text{H}_2\text{(g)} \rightarrow \text{C}_2\text{H}_2\text{(g)}$
- $\text{H}_2\text{O}_2\text{(l)} \rightarrow \text{H}_2\text{(g)} + \text{O}_2\text{(g)}$
- $\text{CH}_2\text{O(l)} + \text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)} + \text{H}_2\text{O(g)}$
- $\text{PbO}_2\text{(s)} \rightarrow \text{Pb(s)} + \text{O}_2\text{(g)}$
- $\text{C}_6\text{H}_{12}\text{(l)} + 9\text{O}_2\text{(g)} \rightarrow 6\text{CO}_2\text{(g)} + 6\text{H}_2\text{O(g)}$
- $3\text{Fe(s)} + 2\text{O}_2\text{(g)} \rightarrow \text{Fe}_3\text{O}_4\text{(s)}$
- $\text{C}_6\text{H}_{12}\text{(l)} + 9\text{O}_2\text{(g)} \rightarrow 6\text{CO}_2\text{(g)} + 6\text{H}_2\text{O(g)}$
- $8\text{SO}_3\text{(g)} \rightarrow \text{S}_8\text{(s)} + 12\text{O}_2\text{(g)}$
- $\text{P}_4\text{(s)} + 6\text{Cl}_2\text{(g)} \rightarrow 4\text{PCl}_3\text{(g)}$
- $8\text{HgS(s)} \rightarrow 8\text{Hg(l)} + \text{S}_8\text{(s)}$
- $\text{C}_7\text{H}_8\text{(l)} + 9\text{O}_2 \rightarrow 7\text{CO}_2\text{(g)} + 4\text{H}_2\text{O(g)}$
- $24\text{C(s)} + 22\text{H}_2\text{(g)} + 11\text{O}_2\text{(g)} \rightarrow 2\text{C}_{12}\text{H}_{22}\text{O}_{11}\text{(s)}$