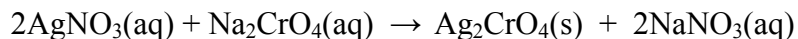


CHAPTER 7	Solution Stoichiometry Problems	BLM 7.2.5
ASSESSMENT		

1. Silver nitrate and sodium chromate solutions react as follows:



(a) How many moles of silver chromate form when 0.200 mol of silver nitrate react?

(b) How many moles of sodium nitrate form when 500 mL of 0.300 mol/L sodium chromate react with silver nitrate?

(c) How many grams of precipitate will form if 200 mL of 0.500 mol/L silver nitrate react with sodium chromate?

2. Write the balanced equation for the reaction of sulfuric acid with a solution of sodium hydroxide.

(a) What volume of 0.320 mol/L sulfuric acid reacts with 47.3 mL of 0.224 mol/L sodium hydroxide?

(b) If 50.0 mL of 0.540 mol/L sodium hydroxide reacts with 85.6 mL of sulfuric acid, calculate the concentration of the sulfuric acid.

CHAPTER 7	Solution Stoichiometry Problems (continued)	BLM 7.2.5
ASSESSMENT		

3. Write the balanced reaction between barium nitrate and sodium carbonate.

(a) How many grams of sodium carbonate would be needed to react with 120 mL of 0.0500 mol/L barium nitrate to entirely remove the barium ions from the solution?

(b) How many grams of precipitate would form if this was done?

4. A student adds 5.77 g of zinc to 100 mL of HCl(aq) . Write the reaction which occurs.

(a) How many moles of HCl(aq) react with the zinc?

(b) What would the concentration of zinc chloride be in the 100 mL solution when all the zinc had reacted?