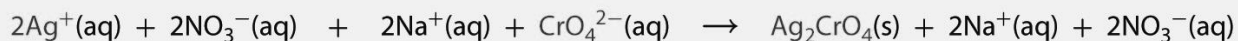
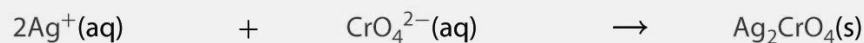
**Non-ionic equation****Total ionic equation****Net ionic equation**

The **complete balanced equation**, or **non-ionic equation**, shows all the reactants and products as if they were intact compounds that had not dissociated in solution.

The **total ionic equation** is a more accurate representation of a reaction in solution, because it shows all the high-solubility ionic compounds dissociated into ions.

The **net ionic equation** shows only the actual chemical change that occurs. To write a net ionic equation, you cancel the spectator ions, or ions which are found in equal quantities among the reactants and the products, from each side of the equation.