

CHAPTER 7	Predicting and Balancing Single and Double Replacement Reactions Answer Key	BLM 7.0.6A
ANSWER KEY		

- $\text{Br}_2(\ell) + 2\text{NaI}(\text{aq}) \rightarrow 2\text{NaBr}(\text{aq}) + \text{I}_2(\text{s})$
- $\text{Ba}(\text{NO}_3)_2(\text{aq}) + \text{K}_2\text{SO}_4(\text{aq}) \rightarrow 2\text{KNO}_3(\text{aq}) + \text{BaSO}_4(\text{s})$
- $\text{CaCl}_2(\text{aq}) + (\text{NH}_4)_2\text{S}(\text{aq}) \rightarrow 2\text{NH}_4\text{Cl}(\text{aq}) + \text{CaS}(\text{s})$
- $\text{Co}(\text{s}) + \text{Cu}(\text{CH}_3\text{COO})_2(\text{aq}) \rightarrow \text{Co}(\text{CH}_3\text{COO})_2(\text{aq}) + \text{Cu}(\text{s})$
- $\text{CH}_3\text{COOH}(\text{aq}) + \text{KOH}(\text{aq}) \rightarrow \text{KCH}_3\text{COO}(\text{aq}) + \text{H}_2\text{O}(\ell)$
- $\text{HgClO}_4(\text{aq}) + \text{NaCH}_3\text{COO}(\text{aq}) \rightarrow \text{HgCH}_3\text{COO}(\text{s}) + \text{NaClO}_4(\text{aq})$
- $\text{Mg}(\text{s}) + 2\text{AgNO}_3(\text{aq}) \rightarrow \text{Mg}(\text{NO}_3)_2(\text{aq}) + 2\text{Ag}(\text{s})$
- $\text{Cd}(\text{ClO}_4)_2(\text{aq}) + \text{Cr}(\text{s}) \rightarrow \text{Cr}(\text{ClO}_4)_2(\text{aq}) + \text{Cd}(\text{s})$
- $\text{AgNO}_3(\text{aq}) + \text{LiCl}(\text{aq}) \rightarrow \text{LiNO}_3(\text{aq}) + \text{AgCl}(\text{s})$
- $2\text{HCN}(\text{aq}) + \text{Ba}(\text{OH})_2(\text{aq}) \rightarrow \text{Ba}(\text{CN})_2(\text{aq}) + 2\text{H}_2\text{O}(\ell)$
- $2\text{Al}(\text{s}) + 3\text{CrSO}_4(\text{aq}) \rightarrow \text{Al}_2(\text{SO}_4)_3(\text{aq}) + 3\text{Cr}(\text{s})$
- $2\text{NaOH}(\text{aq}) + \text{MgI}_2(\text{aq}) \rightarrow \text{Mg}(\text{OH})_2(\text{s}) + 2\text{NaI}(\text{aq})$
- $\text{Cl}_2(\text{aq}) + \text{Na}_2\text{Se}(\text{aq}) \rightarrow 2\text{NaCl}(\text{aq}) + \text{Se}(\text{s})$
- $2\text{HClO}_4(\text{aq}) + \text{Sr}(\text{OH})_2(\text{aq}) \rightarrow \text{Sr}(\text{ClO}_4)_2(\text{aq}) + 2\text{H}_2\text{O}(\ell)$