

CHAPTER 7	Writing Net Ionic Equations Answer Key	BLM 7.1.2A
ANSWER KEY		

- $\text{Fe}^{2+}(\text{aq}) + \text{S}^{2-}(\text{aq}) \rightarrow \text{FeS}(\text{s})$
- $2\text{Ag}^{+}(\text{aq}) + \text{Cu}(\text{s}) \rightarrow \text{Cu}^{2+}(\text{aq}) + 2\text{Ag}(\text{s})$
- $2\text{H}^{+}(\text{aq}) + \text{MgO}(\text{s}) \rightarrow \text{Mg}^{2+}(\text{aq}) + \text{H}_2\text{O}(\ell)$
- $\text{HCN}(\text{aq}) + \text{OH}^{-} \rightarrow \text{H}_2\text{O}(\ell) + \text{CN}^{-}$
- $\text{CO}_2(\text{g}) + 2\text{OH}^{-}(\text{aq}) \rightarrow \text{H}_2\text{O}(\ell) + \text{CO}_3^{2-}(\text{aq})$
- $\text{CO}_3^{2-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow \text{H}_2\text{CO}_3(\text{aq})$
- $3\text{PO}_4^{3-}(\text{aq}) + 2\text{Zn}^{2+}(\text{aq}) \rightarrow \text{Zn}_3(\text{PO}_4)_2(\text{s})$
- $\text{Cd}(\text{s}) + \text{Co}^{2+}(\text{aq}) \rightarrow \text{Cd}^{2+}(\text{aq}) + \text{Co}(\text{s})$
- $\text{CH}_3\text{COOH}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightarrow \text{CH}_3\text{COO}^{-}(\text{aq}) + \text{H}_2\text{O}(\ell)$
- $\text{S}^{2-}(\text{aq}) + 2\text{H}^{+}(\text{aq}) \rightarrow \text{H}_2\text{S}(\text{g})$