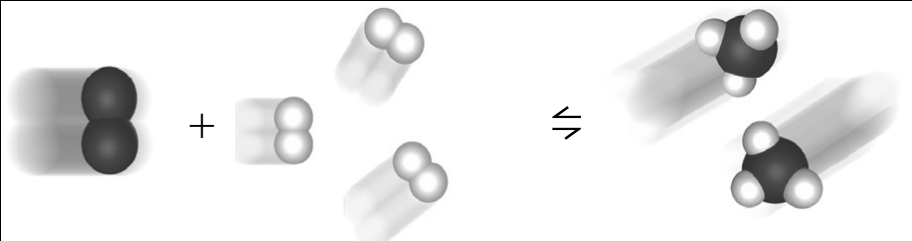


CHAPTER 7	Balanced Chemical Equations	BLM 7.2.1
OVERHEAD		

Balanced equation	$\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$
Number of particles (molecules)	<p>1 molecule <math>\text{N}_2(\text{g}) + 3</math> molecules <math>\text{H}_2(\text{g}) \rightleftharpoons 2</math> molecules <math>\text{NH}_3(\text{g})</math></p> 
Amount (mol)	$1 \text{ mol N}_2(\text{g}) + 3 \text{ mol H}_2(\text{g}) \rightleftharpoons 2 \text{ mol NH}_3(\text{g})$
Mass (g)	$28.02 \text{ g N}_2(\text{g}) + 6.06 \text{ g H}_2(\text{g}) \rightleftharpoons 34.08 \text{ g NH}_3(\text{g})$
Total mass (g)	$34.08 \text{ g reactants} \rightleftharpoons 34.08 \text{ products}$