# Titration Step-By-Step

1. A sample to be analyzed is measured in a graduated or volumetric pipette and delivered to an Erlenmeyer flask.

2. A few drops of the appropriate indicator are added to the sample being analyzed.

3. After taking an initial reading of the volume of titrant in the burette, the titrant is added, slowly, to the sample in the Erlenmeyer flask, while the sample is swirled.
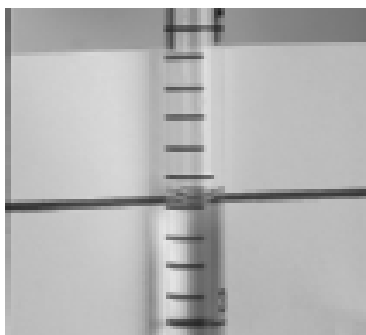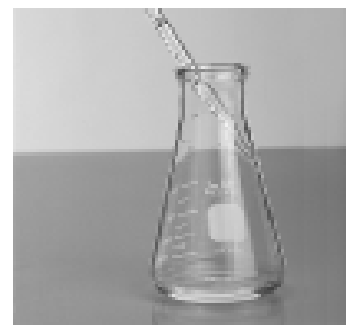
4. The endpoint of the titration occurs when the indicator colour changes dramatically. With a well-chosen indicator, the difference in the volumes of titrant required to reach the endpoint and the equivalence point should differ only by one drop of titrant (approximately 0.05 mL).