

Thermochemical Equations and
Stoichiometry Answer Key

- (a) 1.3×10^3 kJ

(b) 10 kJ

(c) $\Delta H = -1009$ kJ

(d) $\Delta H = -5.2$ kJ
- (a) $\text{CaO(s)} + \text{H}_2\text{O(l)} \rightarrow \text{Ca(OH)}_2 \Delta H = -65.2$ kJ

(b) $\Delta H = -6.08 \times 10^5$ kJ
- (a) 373 kJ

(b) 1.21×10^6 kJ

(c) 2.1 kg