

CHAPTER 5	Oxidation-Reduction Reactions Answer Key	BLM 5.1.6A
ANSWER KEY		

1.
 - a) Compound A is losing an electron; therefore, it is undergoing oxidation (i.e., it is being oxidized).
 - b) Compound B is gaining an electron; therefore, it is undergoing reduction (i.e., it is being reduced).

2.
 - a) True
 - b) False.
A molecule that donates an electron to another molecule is called a reducing agent.
 - c) False.
Compounds contain more energy in their reduced form.
 - d) False.
Oxidation and reduction reactions do not occur independently of each other. When one compound is oxidized (loses an electron) another one must be reduced (it gains the electron)