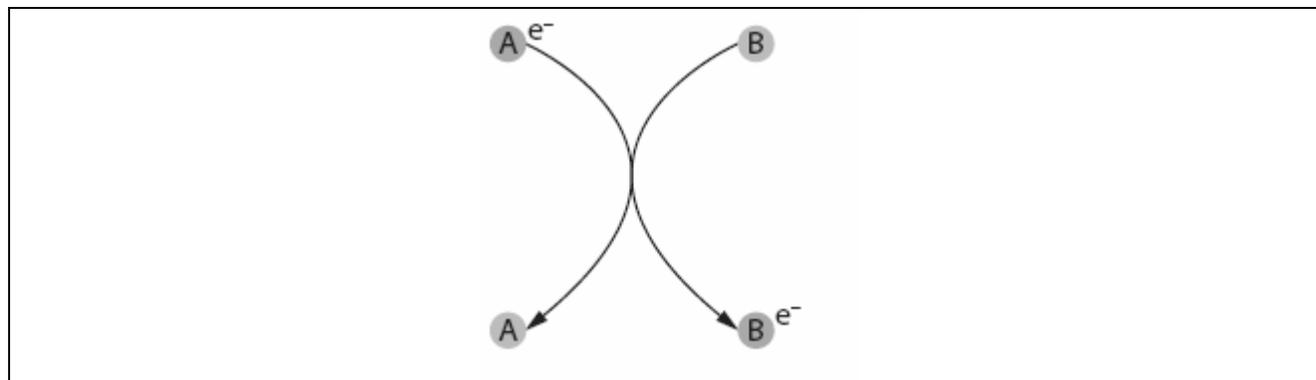


<b>CHAPTER 5</b>	<h1 style="margin: 0;">Oxidation-Reduction Reactions</h1>	<b>BLM 5.1.6</b>
<b>HANDOUT</b>		

The following diagram depicts the transfer of an electron from compound A to compound B. Use this diagram to answer Question 1.



1. a) Is compound A undergoing oxidation or reduction? Why or why not?

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- b) Is compound B undergoing oxidation or reduction? Why or why not?

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2. Indicate whether each of the statements is true or false. If it is false, rewrite the statement to make it true.

TRUE/  
FALSE

- a) An atom or molecule that loses an electron is said to be oxidized.

---

TRUE/  
FALSE

- b) A molecule that donates an electron to another molecule is called an oxidizing agent.

---

TRUE/  
FALSE

- c) Compounds contain more energy in their oxidized form.

---

TRUE/  
FALSE

- d) Oxidations and reductions occur independently of each other and are not linked in any way.

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